

## Bis( $\mu$ -2-methylquinolin-8-olato)- $\kappa^3$ N,O:O; $\kappa^3$ O:N,O-bis[(methanol- $\kappa$ O)-(nitrate- $\kappa^2$ O,O')]lead(II)]

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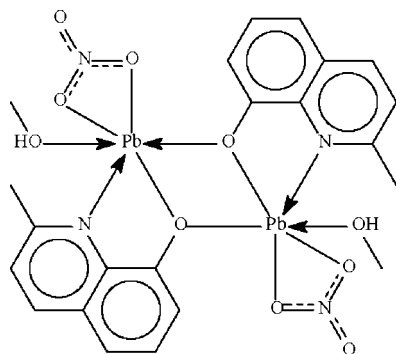
Received 27 January 2009; accepted 29 January 2009

Key indicators: single-crystal X-ray study;  $T = 100$  K; mean  $\sigma(\text{C}-\text{C}) = 0.007$  Å;  $R$  factor = 0.021;  $wR$  factor = 0.058; data-to-parameter ratio = 16.5.

The molecule of the title compound,  $[\text{Pb}_2(\text{C}_{10}\text{H}_8\text{NO})_2(\text{NO}_3)_2(\text{CH}_3\text{OH})_2]$ , lies about a centre of inversion. The Pb atom is chelated by nitrate and substituted quinolin-8-olate anions. The O atom of the quinolin-8-olate also bridges, to confer a six-coordinate status on the metal centre. When a longer Pb...O interaction is considered, the geometry approximates a  $\Psi$ -cube in which one of the sites is occupied by a stereochemically active lone pair.

### Related literature

The 8-hydroxyquinolinolate group engages in  $\mu_3$ -bridging in dinitratohexa(quinolin-8-olato)tetralead(II); see: Zhang *et al.* (2008). It also exhibits this feature in the chain compound, bis(methanol)dinitratodi(quinolin-8-olato)dilead(II); see Shahverdizadeh *et al.* (2008). Both reports comment on lone-pair stereochemistry in this class of lead(II) compounds.



### Experimental

#### Crystal data

$[\text{Pb}_2(\text{C}_{10}\text{H}_8\text{NO})_2(\text{NO}_3)_2(\text{CH}_3\text{O})_2]$	$\gamma = 108.190 (1)^\circ$
$M_r = 918.83$	$V = 630.07 (1) \text{ \AA}^3$
Triclinic, $P\bar{1}$	$Z = 1$
$a = 8.2579 (1) \text{ \AA}$	Mo $K\alpha$ radiation
$b = 8.8052 (1) \text{ \AA}$	$\mu = 13.41 \text{ mm}^{-1}$
$c = 9.6765 (1) \text{ \AA}$	$T = 100 (2) \text{ K}$
$\alpha = 103.976 (1)^\circ$	$0.20 \times 0.15 \times 0.05 \text{ mm}$
$\beta = 98.262 (1)^\circ$	

#### Data collection

Bruker SMART APEX diffractometer	5958 measured reflections
Absorption correction: multi-scan (SADABS; Sheldrick, 1996)	2872 independent reflections
$T_{\min} = 0.175$ , $T_{\max} = 0.554$ (expected range = 0.162–0.512)	2754 reflections with $I > 2\sigma(I)$
	$R_{\text{int}} = 0.019$

#### Refinement

$R[F^2 > 2\sigma(F^2)] = 0.021$	H atoms treated by a mixture of independent and constrained refinement
$wR(F^2) = 0.058$	$\Delta\rho_{\text{max}} = 1.41 \text{ e \AA}^{-3}$
$S = 1.08$	$\Delta\rho_{\text{min}} = -2.12 \text{ e \AA}^{-3}$
2872 reflections	
174 parameters	
1 restraint	

Data collection: APEX2 (Bruker, 2008); cell refinement: SAINT (Bruker, 2008); data reduction: SAINT; program(s) used to solve structure: SHELXS97 (Sheldrick, 2008); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: X-SEED (Barbour, 2001); software used to prepare material for publication: publCIF (Westrip, 2009).

The authors thank Shahid Beheshti University and the University of Malaya for supporting this study.

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: TK2363).

### References

- Barbour, L. J. (2001). *J. Supramol. Chem.* **1**, 189–191.  
 Bruker (2008). APEX2 and SAINT. Bruker AXS Inc., Madison, Wisconsin, USA.  
 Shahverdizadeh, G. H., Soudi, A. A., Morsali, A. & Retailleau, P. (2008). *Inorg. Chim. Acta*, **361**, 1875–1884.  
 Sheldrick, G. M. (1996). SADABS. University of Göttingen, Germany.  
 Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.  
 Westrip, S. P. (2009). publCIF. In preparation.  
 Zhang, W.-Z., Wei, D.-Z., Che, X.-F., Gao, E.-J., Wang, K.-H., Yin, H.-X. & Gu, X.-G. (2008). *Chin. J. Struct. Chem.* **27**, 287–292.

**supplementary materials**

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**Bis( $\mu$ -2-methylquinolin-8-olato)- $\kappa^3N,O:O;\kappa^3O:N,O$ -bis[(methanol- $\kappa O$ )(nitrate- $\kappa^2O,O'$ )lead(II)]**

**G. Mohammadnezhad Sh., M. M. Amini and S. W. Ng**

**Experimental**

Lead nitrate (0.33 g, 1 mmol) and 2-methyl-8-hydroxyquinoline (0.32 g, 2 mmol) were loaded into a convection tube; the tube was filled with dry methanol and kept at 333 K. Crystals were collected from the side arm after 3 d.

**Refinement**

Carbon-bound H atoms were placed in calculated positions (C—H 0.93 to 0.98 Å) and were included in the refinement in the riding model approximation, with  $U(H)$  set to 1.2 to 1.5 $U(C)$ . The methanol H atom was located in a difference Fourier map, and was refined with a distance restraint of O—H 0.84 (1) Å; its temperature factor was freely refined.

The crystal diffracted strongly owing to the extremely heavy metal atom; however, its presence introduced severe absorption problems that could not be corrected analytically as the crystal did not have regular faces. Although a sphere of reflections was measured, multi-scan treatment only marginally improved the quality. The final difference Fourier map had large peaks/deep holes near the Pb atom.

**Figures**

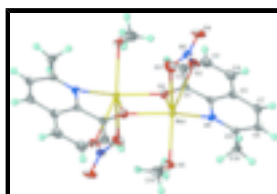


Fig. 1. Thermal ellipsoid plot (Barbour, 2001) of  $Pb_2(NO_3)_2(CH_4O)_2(C_{10}H_8NO)_2$ ; ellipsoids are drawn at the 70% probability level and H atoms of arbitrary radius. The unlabelled atoms are related by 1-x, 2-y, 1-z.

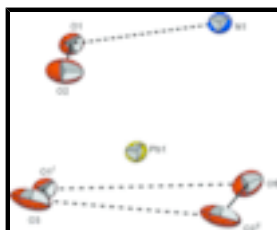


Fig. 2. Detail of the environment of the Pb atom. Symmetry codes: (i) 1 - x, 2 - y, 1 - z; (ii) 1 - x, 1 - y, 1 - z.

**Bis( $\mu$ -2-methylquinolin-8-olato)- $\kappa^3N,O:O;\kappa^3O:N,O$ - bis[(methanol- $\kappa O$ )(nitrate- $\kappa^2O,O'$ )lead(II)]**

*Crystal data*

[ $Pb_2(C_{10}H_8NO)_2(NO_3)_2(CH_4O)_2$ ]

$M_r = 918.83$

Triclinic,  $P\bar{1}$

$Z = 1$

$F_{000} = 428$

$D_x = 2.422 \text{ Mg m}^{-3}$

# supplementary materials

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Hall symbol: -P 1

$a = 8.2579 (1) \text{ \AA}$

$b = 8.8052 (1) \text{ \AA}$

$c = 9.6765 (1) \text{ \AA}$

$\alpha = 103.976 (1)^\circ$

$\beta = 98.262 (1)^\circ$

$\gamma = 108.190 (1)^\circ$

$V = 630.07 (1) \text{ \AA}^3$

Mo  $K\alpha$  radiation

$\lambda = 0.71073 \text{ \AA}$

Cell parameters from 5067 reflections

$\theta = 2.2\text{--}28.3^\circ$

$\mu = 13.41 \text{ mm}^{-1}$

$T = 100 \text{ K}$

Block, yellow

$0.20 \times 0.15 \times 0.05 \text{ mm}$

## Data collection

Bruker SMART APEX  
diffractometer

Radiation source: fine-focus sealed tube

Monochromator: graphite

$T = 100 \text{ K}$

$\omega$  scans

Absorption correction: multi-scan  
(SADABS; Sheldrick, 1996)

$T_{\min} = 0.175$ ,  $T_{\max} = 0.554$

5958 measured reflections

2872 independent reflections

2754 reflections with  $I > 2\sigma(I)$

$R_{\text{int}} = 0.019$

$\theta_{\max} = 27.5^\circ$

$\theta_{\min} = 2.2^\circ$

$h = -10 \rightarrow 10$

$k = -11 \rightarrow 11$

$l = -12 \rightarrow 12$

## Refinement

Refinement on  $F^2$

Least-squares matrix: full

$R[F^2 > 2\sigma(F^2)] = 0.021$

$wR(F^2) = 0.058$

$S = 1.08$

2872 reflections

174 parameters

1 restraint

Primary atom site location: structure-invariant direct  
methods

Secondary atom site location: difference Fourier map

Hydrogen site location: inferred from neighbouring  
sites

H atoms treated by a mixture of  
independent and constrained refinement

$$w = 1/[\sigma^2(F_o^2) + (0.0341P)^2 + 1.3909P]$$

where  $P = (F_o^2 + 2F_c^2)/3$

$(\Delta/\sigma)_{\max} = 0.001$

$\Delta\rho_{\max} = 1.41 \text{ e \AA}^{-3}$

$\Delta\rho_{\min} = -2.12 \text{ e \AA}^{-3}$

Extinction correction: none

## Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )

	<i>x</i>	<i>y</i>	<i>z</i>	$U_{\text{iso}}^*/U_{\text{eq}}$
Pb1	0.403585 (18)	0.747859 (17)	0.434423 (15)	0.01020 (6)
O1	0.5672 (4)	0.9839 (4)	0.3886 (3)	0.0139 (6)
O2	0.7047 (5)	0.7066 (4)	0.4111 (4)	0.0202 (7)
O3	0.6609 (5)	0.6459 (5)	0.6104 (4)	0.0263 (8)
O4	0.8652 (5)	0.5918 (5)	0.5137 (4)	0.0261 (8)
O5	0.0903 (4)	0.7403 (4)	0.3446 (4)	0.0175 (6)

H5	0.006 (6)	0.691 (7)	0.376 (7)	0.028 (17)*
N1	0.3858 (5)	0.7002 (4)	0.1665 (4)	0.0114 (7)
N2	0.7441 (5)	0.6484 (5)	0.5124 (4)	0.0146 (7)
C1	0.6165 (6)	0.9696 (5)	0.2615 (5)	0.0123 (8)
C2	0.7547 (6)	1.0900 (6)	0.2411 (5)	0.0169 (9)
H2	0.8219	1.1885	0.3203	0.020*
C3	0.7985 (6)	1.0691 (6)	0.1028 (5)	0.0205 (9)
H3	0.8963	1.1529	0.0913	0.025*
C4	0.7024 (6)	0.9308 (6)	-0.0140 (5)	0.0199 (9)
H4	0.7315	0.9203	-0.1065	0.024*
C5	0.5592 (6)	0.8028 (5)	0.0032 (5)	0.0141 (8)
C6	0.5180 (5)	0.8209 (5)	0.1415 (4)	0.0115 (8)
C7	0.4520 (6)	0.6577 (6)	-0.1116 (5)	0.0167 (9)
H7	0.4731	0.6417	-0.2070	0.020*
C8	0.3161 (6)	0.5389 (6)	-0.0852 (5)	0.0165 (9)
H8	0.2419	0.4414	-0.1626	0.020*
C9	0.2879 (5)	0.5628 (5)	0.0573 (5)	0.0121 (8)
C10	0.1480 (5)	0.4306 (4)	0.0890 (5)	0.0160 (9)
H10A	0.2007	0.3646	0.1358	0.024*
H10B	0.0632	0.3567	-0.0028	0.024*
H10C	0.0879	0.4834	0.1549	0.024*
C11	0.0687 (6)	0.8890 (4)	0.3246 (5)	0.0257 (11)
H11A	0.0928	0.9717	0.4205	0.039*
H11B	-0.0520	0.8616	0.2712	0.039*
H11C	0.1504	0.9358	0.2684	0.039*

Atomic displacement parameters ( $\text{\AA}^2$ )

	$U^{11}$	$U^{22}$	$U^{33}$	$U^{12}$	$U^{13}$	$U^{23}$
Pb1	0.01050 (9)	0.00980 (9)	0.01087 (9)	0.00331 (6)	0.00258 (6)	0.00458 (6)
O1	0.0147 (15)	0.0116 (13)	0.0134 (14)	0.0013 (12)	0.0045 (12)	0.0043 (11)
O2	0.0235 (18)	0.0268 (17)	0.0212 (16)	0.0148 (14)	0.0101 (14)	0.0161 (14)
O3	0.0270 (19)	0.046 (2)	0.0163 (16)	0.0181 (17)	0.0125 (14)	0.0177 (16)
O4	0.0162 (17)	0.0333 (19)	0.039 (2)	0.0136 (15)	0.0108 (15)	0.0205 (17)
O5	0.0129 (16)	0.0205 (16)	0.0227 (16)	0.0073 (13)	0.0072 (13)	0.0098 (13)
N1	0.0117 (17)	0.0121 (16)	0.0120 (16)	0.0041 (14)	0.0028 (13)	0.0067 (13)
N2	0.0126 (18)	0.0164 (17)	0.0159 (17)	0.0051 (14)	0.0036 (14)	0.0068 (14)
C1	0.014 (2)	0.0129 (18)	0.0128 (19)	0.0072 (16)	0.0030 (16)	0.0055 (16)
C2	0.017 (2)	0.015 (2)	0.017 (2)	0.0048 (17)	0.0033 (17)	0.0049 (17)
C3	0.018 (2)	0.019 (2)	0.027 (2)	0.0039 (18)	0.0113 (19)	0.0115 (19)
C4	0.021 (2)	0.022 (2)	0.022 (2)	0.0080 (19)	0.0124 (19)	0.0110 (19)
C5	0.015 (2)	0.017 (2)	0.0138 (19)	0.0064 (17)	0.0063 (16)	0.0096 (17)
C6	0.0095 (19)	0.0148 (19)	0.0134 (19)	0.0071 (16)	0.0022 (15)	0.0067 (16)
C7	0.018 (2)	0.021 (2)	0.014 (2)	0.0101 (18)	0.0053 (17)	0.0049 (17)
C8	0.014 (2)	0.017 (2)	0.017 (2)	0.0061 (17)	0.0020 (16)	0.0030 (17)
C9	0.0088 (19)	0.0137 (19)	0.0140 (19)	0.0047 (16)	0.0011 (15)	0.0045 (16)
C10	0.011 (2)	0.0125 (19)	0.020 (2)	-0.0014 (16)	0.0022 (16)	0.0038 (16)
C11	0.022 (3)	0.024 (2)	0.036 (3)	0.010 (2)	0.008 (2)	0.014 (2)

## supplementary materials

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### Geometric parameters ( $\text{\AA}$ , $^\circ$ )

Pb1—O1	2.281 (3)	C2—C3	1.421 (6)
Pb1—O1 <sup>i</sup>	2.478 (3)	C2—H2	0.9500
Pb1—N1	2.499 (3)	C3—C4	1.365 (7)
Pb1—O5	2.583 (3)	C3—H3	0.9500
Pb1—O2	2.655 (3)	C4—C5	1.418 (6)
Pb1—O3	3.019 (4)	C4—H4	0.9500
Pb1—O3 <sup>ii</sup>	3.248 (4)	C5—C7	1.406 (6)
Pb1—O4 <sup>ii</sup>	3.320 (4)	C5—C6	1.412 (6)
O1—C1	1.341 (5)	C7—C8	1.378 (6)
O1—Pb1 <sup>i</sup>	2.478 (3)	C7—H7	0.9500
O2—N2	1.259 (5)	C8—C9	1.408 (6)
O3—N2	1.248 (5)	C8—H8	0.9500
O4—N2	1.248 (5)	C9—C10	1.490 (5)
O5—C11	1.430 (5)	C10—H10A	0.9800
O5—H5	0.838 (10)	C10—H10B	0.9800
N1—C9	1.326 (5)	C10—H10C	0.9800
N1—C6	1.362 (5)	C11—H11A	0.9800
C1—C2	1.369 (6)	C11—H11B	0.9800
C1—C6	1.433 (6)	C11—H11C	0.9800
O1—Pb1—O1 <sup>i</sup>	64.88 (12)	O4—N2—O2	119.8 (4)
O1—Pb1—N1	68.38 (11)	O1—C1—C2	123.2 (4)
O1 <sup>i</sup> —Pb1—N1	124.94 (11)	O1—C1—C6	117.8 (4)
O1—Pb1—O5	100.85 (11)	C2—C1—C6	119.0 (4)
O1 <sup>i</sup> —Pb1—O5	82.53 (10)	C1—C2—C3	120.6 (4)
N1—Pb1—O5	79.21 (11)	C1—C2—H2	119.7
O1—Pb1—O2	75.46 (11)	C3—C2—H2	119.7
O1 <sup>i</sup> —Pb1—O2	114.95 (11)	C4—C3—C2	121.2 (4)
N1—Pb1—O2	78.25 (11)	C4—C3—H3	119.4
O5—Pb1—O2	156.87 (11)	C2—C3—H3	119.4
O1—Pb1—O3	105.98 (11)	C3—C4—C5	119.8 (4)
O1 <sup>i</sup> —Pb1—O3	100.58 (10)	C3—C4—H4	120.1
N1—Pb1—O3	119.16 (10)	C5—C4—H4	120.1
O5—Pb1—O3	151.65 (10)	C7—C5—C6	117.2 (4)
O2—Pb1—O3	44.42 (9)	C7—C5—C4	123.5 (4)
O1—Pb1—O3 <sup>ii</sup>	144.45 (10)	C6—C5—C4	119.3 (4)
O1 <sup>i</sup> —Pb1—O3 <sup>ii</sup>	144.84 (9)	N1—C6—C5	122.2 (4)
N1—Pb1—O3 <sup>ii</sup>	89.96 (10)	N1—C6—C1	117.7 (4)
O5—Pb1—O3 <sup>ii</sup>	102.30 (10)	C5—C6—C1	120.2 (4)
O2—Pb1—O3 <sup>ii</sup>	72.63 (10)	C8—C7—C5	119.8 (4)
O3—Pb1—O3 <sup>ii</sup>	59.42 (12)	C8—C7—H7	120.1
O1—Pb1—O4 <sup>ii</sup>	175.03 (10)	C5—C7—H7	120.1
O1 <sup>i</sup> —Pb1—O4 <sup>ii</sup>	114.40 (10)	C7—C8—C9	119.6 (4)

N1—Pb1—O4 <sup>ii</sup>	109.62 (11)	C7—C8—H8	120.2
O5—Pb1—O4 <sup>ii</sup>	74.21 (9)	C9—C8—H8	120.2
O2—Pb1—O4 <sup>ii</sup>	108.83 (9)	N1—C9—C8	121.5 (4)
O3—Pb1—O4 <sup>ii</sup>	78.99 (10)	N1—C9—C10	118.5 (4)
O3 <sup>ii</sup> —Pb1—O4 <sup>ii</sup>	38.47 (9)	C8—C9—C10	120.0 (4)
C1—O1—Pb1	119.2 (2)	C9—C10—H10A	109.5
C1—O1—Pb1 <sup>i</sup>	124.9 (2)	C9—C10—H10B	109.5
Pb1—O1—Pb1 <sup>i</sup>	115.12 (12)	H10A—C10—H10B	109.5
N2—O2—Pb1	106.1 (2)	C9—C10—H10C	109.5
N2—O3—Pb1	88.6 (2)	H10A—C10—H10C	109.5
C11—O5—Pb1	118.7 (2)	H10B—C10—H10C	109.5
C11—O5—H5	108 (4)	O5—C11—H11A	109.5
Pb1—O5—H5	122 (4)	O5—C11—H11B	109.5
C9—N1—C6	119.7 (4)	H11A—C11—H11B	109.5
C9—N1—Pb1	127.4 (3)	O5—C11—H11C	109.5
C6—N1—Pb1	111.6 (3)	H11A—C11—H11C	109.5
O3—N2—O4	120.2 (4)	H11B—C11—H11C	109.5
O3—N2—O2	119.9 (4)		
O1 <sup>i</sup> —Pb1—O1—C1	-170.3 (4)	O1—Pb1—N1—C6	17.3 (3)
N1—Pb1—O1—C1	-20.3 (3)	O1 <sup>i</sup> —Pb1—N1—C6	50.8 (3)
O5—Pb1—O1—C1	-93.9 (3)	O5—Pb1—N1—C6	123.7 (3)
O2—Pb1—O1—C1	62.6 (3)	O2—Pb1—N1—C6	-61.5 (3)
O3—Pb1—O1—C1	95.3 (3)	O3—Pb1—N1—C6	-79.4 (3)
O3 <sup>ii</sup> —Pb1—O1—C1	36.0 (4)	O3 <sup>ii</sup> —Pb1—N1—C6	-133.8 (3)
O4 <sup>ii</sup> —Pb1—O1—C1	-87.4 (12)	O4 <sup>ii</sup> —Pb1—N1—C6	-167.5 (3)
O1 <sup>i</sup> —Pb1—O1—Pb1 <sup>i</sup>	0.0	Pb1—O3—N2—O4	-170.3 (4)
N1—Pb1—O1—Pb1 <sup>i</sup>	150.03 (17)	Pb1—O3—N2—O2	9.0 (4)
O5—Pb1—O1—Pb1 <sup>i</sup>	76.34 (14)	Pb1—O2—N2—O3	-10.7 (5)
O2—Pb1—O1—Pb1 <sup>i</sup>	-127.07 (15)	Pb1—O2—N2—O4	168.6 (3)
O3—Pb1—O1—Pb1 <sup>i</sup>	-94.40 (14)	Pb1—O1—C1—C2	-159.2 (3)
O3 <sup>ii</sup> —Pb1—O1—Pb1 <sup>i</sup>	-153.76 (12)	Pb1 <sup>i</sup> —O1—C1—C2	31.5 (6)
O4 <sup>ii</sup> —Pb1—O1—Pb1 <sup>i</sup>	82.9 (12)	Pb1—O1—C1—C6	21.2 (5)
O1—Pb1—O2—N2	137.6 (3)	Pb1 <sup>i</sup> —O1—C1—C6	-148.1 (3)
O1 <sup>i</sup> —Pb1—O2—N2	84.7 (3)	O1—C1—C2—C3	-178.8 (4)
N1—Pb1—O2—N2	-152.0 (3)	C6—C1—C2—C3	0.8 (7)
O5—Pb1—O2—N2	-138.9 (3)	C1—C2—C3—C4	1.4 (7)
O3—Pb1—O2—N2	5.4 (2)	C2—C3—C4—C5	-1.8 (7)
O3 <sup>ii</sup> —Pb1—O2—N2	-58.3 (3)	C3—C4—C5—C7	179.0 (5)
O4 <sup>ii</sup> —Pb1—O2—N2	-45.1 (3)	C3—C4—C5—C6	0.1 (7)
O1—Pb1—O3—N2	-53.6 (3)	C9—N1—C6—C5	-2.1 (6)
O1 <sup>i</sup> —Pb1—O3—N2	-120.3 (3)	Pb1—N1—C6—C5	166.0 (3)
N1—Pb1—O3—N2	20.2 (3)	C9—N1—C6—C1	178.0 (4)
O5—Pb1—O3—N2	145.9 (3)	Pb1—N1—C6—C1	-13.9 (4)
O2—Pb1—O3—N2	-5.3 (2)	C7—C5—C6—N1	3.2 (6)

## supplementary materials

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O3 <sup>ii</sup> —Pb1—O3—N2	90.9 (3)	C4—C5—C6—N1	-177.8 (4)
O4 <sup>ii</sup> —Pb1—O3—N2	126.7 (3)	C7—C5—C6—C1	-176.9 (4)
O1—Pb1—O5—C11	-16.8 (3)	C4—C5—C6—C1	2.0 (6)
O1 <sup>i</sup> —Pb1—O5—C11	45.7 (3)	O1—C1—C6—N1	-3.0 (6)
N1—Pb1—O5—C11	-82.1 (3)	C2—C1—C6—N1	177.4 (4)
O2—Pb1—O5—C11	-95.2 (4)	O1—C1—C6—C5	177.1 (4)
O3—Pb1—O5—C11	144.2 (3)	C2—C1—C6—C5	-2.5 (6)
O3 <sup>ii</sup> —Pb1—O5—C11	-169.7 (3)	C6—C5—C7—C8	-1.5 (6)
O4 <sup>ii</sup> —Pb1—O5—C11	163.8 (3)	C4—C5—C7—C8	179.6 (4)
O1—Pb1—N1—C9	-175.7 (4)	C5—C7—C8—C9	-1.0 (7)
O1 <sup>i</sup> —Pb1—N1—C9	-142.2 (3)	C6—N1—C9—C8	-0.7 (6)
O5—Pb1—N1—C9	-69.4 (3)	Pb1—N1—C9—C8	-166.7 (3)
O2—Pb1—N1—C9	105.4 (4)	C6—N1—C9—C10	178.1 (4)
O3—Pb1—N1—C9	87.5 (4)	Pb1—N1—C9—C10	12.1 (5)
O3 <sup>ii</sup> —Pb1—N1—C9	33.2 (4)	C7—C8—C9—N1	2.3 (7)
O4 <sup>ii</sup> —Pb1—N1—C9	-0.6 (4)	C7—C8—C9—C10	-176.5 (4)

Symmetry codes: (i)  $-x+1, -y+2, -z+1$ ; (ii)  $-x+1, -y+1, -z+1$ .

### Hydrogen-bond geometry ( $\text{\AA}$ , $^\circ$ )

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
O5—H5 <sup>iii</sup> —O4 <sup>iii</sup>	0.84 (1)	2.06 (2)	2.869 (5)	161 (6)

Symmetry codes: (iii)  $x-1, y, z$ .



Fig. 1

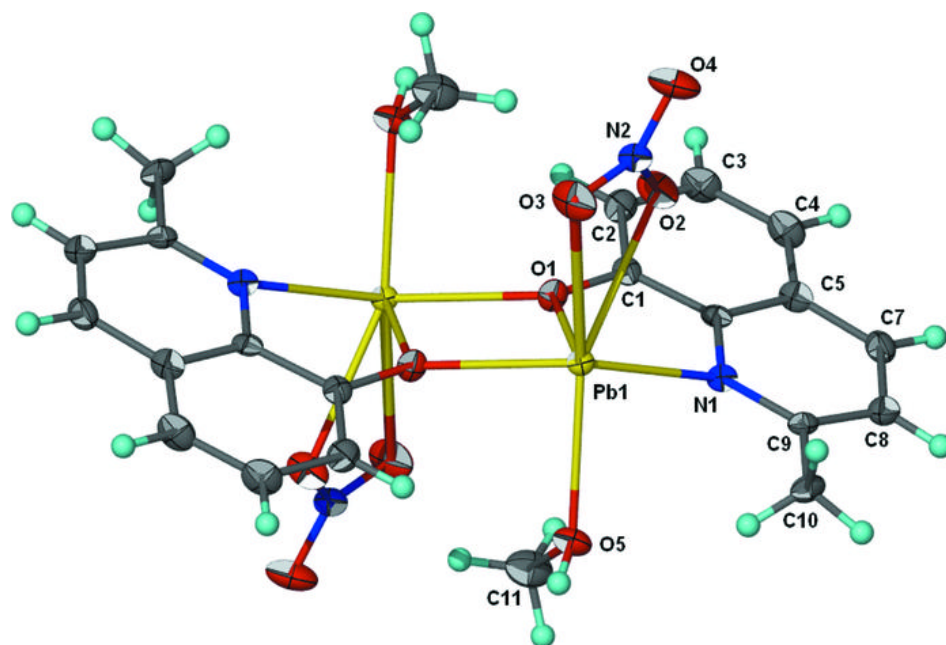


Fig. 2

